

# Controlling the decoration of the reduced graphene oxide surface with pyrene functionalized gold nanoparticles

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## Abstract

We exploited a non-covalent approach based on  $\pi$ -stacking interactions to address the formation of hybrids between pyrene-functionalized gold nanoparticles (PyAuNPs) and reduced graphene oxide (RGO), in which the former are distributed homogeneously on the surface of the latter with a high degree of coverage. We used water soluble PyAuNPs of two different average dimensions, namely 2 and 8 nm, in which the pendant pyrene moieties were introduced within a mixed monolayer with a choline derivative. The combination with RGO originates highly insoluble materials, in which microscopy evidences a complete adhesion of the PyAuNPs onto the carbon nanomaterial layers in a highly homogeneous fashion, with no traces of free particles, confirming the high affinity between pyrene-functionalized species and conjugated carbon nanostructure surfaces.

## Keywords

Reduced graphene oxide, gold nanoparticles, functional nanohybrids

## 1. Introduction

The combination of graphene based materials<sup>[1]</sup> (GBMs) with gold nanoparticles<sup>[2]</sup> (AuNPs) is extremely promising for the preparation of functional hybrid materials for catalysis,<sup>[3, 4]</sup> photonics,<sup>[5]</sup> electronics,<sup>[6, 7]</sup> sensing<sup>[8-12]</sup> and various kind of biomedical<sup>[8, 13]</sup> applications. Both materials have in fact interesting properties *per se*: GBMs demonstrate outstanding electrical<sup>[14]</sup> and thermal<sup>[15]</sup> conductivities coupled to mechanical strength<sup>[16]</sup> and high optical transparency,<sup>[17]</sup> whereas AuNPs have tunable absorption properties depending on the size and shape, can induce metal-enhanced fluorescence in fluorophores<sup>[18]</sup> and can be stabilized with a large plethora of organic moieties exploiting alkanethiols-gold bonds, thus also introducing further selective functions.<sup>[19-21]</sup>

Among GBMs, reduced graphene oxide (RGO) emerges as a good compromise between materials with impressive electronic properties, such as those which can be prepared through chemical vapor deposition of suitable precursors on top of a catalytic surface,<sup>[22]</sup> and materials that can be produced

1 on a large scale with limited costs.<sup>[23]</sup> It is in fact obtained via a two-step procedure starting from  
2 largely available natural graphite.<sup>[24]</sup>

3 Strategies to produce RGO-AuNPs composites normally involve the *in situ* reduction of a suitable  
4 gold precursor in the presence of the GBM, or of its oxidized precursor, graphene oxide (GO), which  
5 is also reduced to RGO during the process.<sup>[7, 25-27]</sup> Defects and oxygen residues on the surface of RGO  
6 can act as nucleating sites for the growth of the metal nanoparticles. This approach can be highly  
7 efficient and easy to perform, but also very difficult to control in terms of achieving uniform sizes for  
8 the *in situ* prepared AuNPs as well as in the homogeneity of their distribution onto the RGO substrate.  
9 *Ex situ* assembly of previously prepared AuNPs on RGO can be exploited as an alternative strategy,  
10 allowing to start from nanoparticle species of which it is easier to control both the shape and size  
11 distribution and which can feature particular functionalities aimed at establishing high affinity  
12 interactions with the GBM. Also in this case, the presence of oxygenated functional groups on RGO  
13 can furnish a valuable tool to build up covalent bonds with AuNPs, if the proper reactive functional  
14 groups are introduced on these last ones. Alternatively carbon-carbon double bonds on RGO can be  
15 involved in other types of reactions, such as electrophilic additions or cycloadditions.<sup>[28, 29]</sup> A milder  
16 alternative to the use of covalent chemistry involves the so-called non-covalent functionalization of  
17 RGO by means of  $\pi$ -conjugated molecular systems, mainly pyrene, anthracene and naphthalene, but  
18 also more complex ones containing heteroatoms.<sup>[30]</sup> As stated before, being the chemistry of AuNPs  
19 so versatile in terms of chemical functionalities which can be appended to the gold surface through  
20 Au-S bonds, making use of selected alkanethiols, one can easily design and synthesize AuNPs  
21 featuring pendant  $\pi$ -conjugated molecules. Brunetti et al. made use of AuNPs capped with an  
22 alkanethiol functionalized 9,10-di(1,3-dithiole-2-ylidene)-9,10-dihydroanthracene (exTTF) to  
23 prepare non-covalent nanohybrids with exfoliated graphene.<sup>[31]</sup> They demonstrated that analogous  
24 AuNPs lacking the exTTF moiety were not interacting with graphene, thus highlighting the key role  
25 played by  $\pi$ - $\pi$  stacking interactions in driving the formation of the nanohybrid. Furthermore, the  
26 exTTF functionalized AuNPs were shown to be effective in binding to graphene because of the  
27 presence of multivalent interactions, whereas by employing simple exTTF molecules no binding was  
28 achieved. Notwithstanding that well-defined graphene-AuNPs species were obtained following this  
29 approach, no homogeneous dispersion of the particles on the graphene surface was evident from  
30 transmission electron microscopy (TEM) images, but rather islands of aggregated AuNPs separated  
31 by areas where no graphene decoration was present.

32 We report here a similar approach for the preparation of RGO-AuNPs nanohybrids, which we have  
33 previously employed for the preparation of analogous species with single walled carbon nanotubes  
34 (SWCNTs).<sup>[32]</sup> The strategy involves the synthesis of AuNPs capped with a mixed monolayer of

1 choline and pyrene derivatives (PyAuNPs), being the latter useful for the interaction with the carbon  
2 nanostructure surface. In the case of SWCNTs, we demonstrated that, by employing 2 nm PyAuNPs,  
3 water soluble nanohybrids were generated, thus being able to make use of the water affinity of choline  
4 moieties to extract hydrophobic pristine nanotubes in a polar environment.<sup>[32]</sup> In the present case we  
5 obtain insoluble RGO-AuNPs nanohybrids, characterized by a very homogeneous distribution of  
6 AuNPs on the surface of RGO, stating the validity of our mild non-covalent method to obtain well-  
7 defined composite materials.  
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## 9 **2. Experimental section**

10 **General:** Solvents were purified by standard methods. All commercially reagents were purchased  
11 from Sigma-Aldrich, unless not specified, and used as received. Thin layer chromatography analyses  
12 were performed using Merck 60 F<sub>254</sub> precoated silica gel glass plates. Column chromatography was  
13 carried out on Macherey-Nagel silica gel 60 (70-230 mesh). NMR spectra were recorded using a  
14 Bruker AV III 500 spectrometer operating at 500 MHz for <sup>1</sup>H, 125.8 MHz for <sup>13</sup>C. Chemical shifts  
15 are reported relative to internal Me<sub>4</sub>Si. Multiplicity is given as follow: s = singlet, d = doublet, t =  
16 triplet, q = quartet, qn = quintet, m = multiplet, br = broad peak. HRMS mass spectra were obtained  
17 with a Mariner Applied Biosystem (API-TOF) mass spectrometer (MeOH, 0.5% formic acid). TEM  
18 images were recorded on a Jeol 300 PX electron microscope. One drop of sample was placed on the  
19 sample grid and the solvent was allowed to evaporate. UV-Visible spectra were recorded on a Cary  
20 50 spectrophotometer equipped with thermostated multiple cell holders. The Raman spectra were  
21 acquired with an Xplora spectrometer from Horiba with an excitation line at 632 nm. Excitation  
22 photoluminescence spectroscopy was performed with a Nanolog spectrofluorometer from Horiba,  
23 equipped with a Xenon lamp and a liquid-nitrogen cooled InGaAs detector.  
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### 25 **2.1. Synthesis of thiol 1 and 2**

26 Thiol 1 and 2 were prepared according to previously reported procedure.<sup>[32, 33]</sup>

### 27 **2.2 Synthesis and characterization of monolayer protected gold nanoparticles (T1-NP-a, b)**

28 **NP-a** (2 nm) monolayer protected gold nanoparticles were prepared modifying a previously reported  
29 two-step procedure.<sup>[34]</sup> A solution of H<sub>2</sub>AuCl<sub>4</sub>·3H<sub>2</sub>O (50 mg, 0.127 mmol, 1 equiv) in water (2 mL)  
30 was extracted with a solution of tetraoctylammonium bromide (0.175 g, 0.318 mmol, 2.5 equiv) in  
31 N<sub>2</sub> purged toluene (125 mL). To the resulting reddish-orange organic solution dioctylamine (0.613 g,  
32 2.539 mmol, 20 equiv) was added (the amount of dioctylamine was calculated in order to obtain 2  
33 nm nanoparticles). The mixture was vigorously stirred under N<sub>2</sub> for 1,5 hours. During this period of  
34 time the color of the mixture faded. Then a NaBH<sub>4</sub> solution (48.0 mg, 1.269 mmol, 10 equiv) in H<sub>2</sub>O  
35 (1 mL) was then rapidly added. The color of the solution turned rapidly to black and after 1.5 hours  
36 of stirring at 0°C, the aqueous layer was removed. To the obtained nanoparticle solution, thiol 2  
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(0.254 mmol, 2 equiv) dissolved in 3 mL of MeOH was rapidly added. The reaction mixture was stirred overnight. Then the mixture was centrifuged and the precipitated NPs were dissolved in the minimum amount of MeOH, sonicated and then precipitated with EtOAc (5 times); the resulting NPs were finally purified by gel permeation chromatography with Sephadex LH-20.

**NP-b** (8 nm) monolayer protected gold nanoparticles were prepared modifying a reported procedure.<sup>[35]</sup> Typically, to 5.6 ml of water, sequentially and under vigorous stirring, 1.6 ml of sodium citrate (510 mM), 250  $\mu$ l of silver nitrate (10 mM) and 500  $\mu$ l of tetrachloroauric acid (250 mM) were added and stirred for 5 minutes. During this time, the solution changed from initial yellow color to green, grey and finally black. After the incubation time, the solution was quickly added to 117 ml of boiling water and heated under reflux for 1 hour, becoming wine-red after a few seconds. The citrate-capped nanoparticles solution obtained was then allowed to cool down to room temperature, then thiol **2** was added in 1-2 ml of isopropanol and left stirring overnight. Nanoparticles were purified by using 15 ml Amicon<sup>®</sup> Ultra filters of 100 kDa molecular weight cutoff centrifuged for 2.5 minutes at 2000 rpm or 10 kDa at 5000 rpm for 6 minutes. Prior to use, filters were prewashed 5 times with 1:1 H<sub>2</sub>O:MeOH, then nanoparticles washed 5 times with 1:1 H<sub>2</sub>O:MeOH and 5 times with pure water. Typical concentrated volumes after centrifugation for 15, 4 and 0.5 ml filtering units were 1.5, 0.5 and 0.1 ml respectively. Purified nanoparticles aqueous solution were frozen with dry ice and freeze-dried overnight on a 5Pascal 105PDGT lyophilizer equipped with and Edwards XDS 10 Pump (P  $\approx$  0.2 mbar, T = -50 °C). Typically, an amount between 22 and 26 mg of functionalized nanoparticles was obtained, corresponding to *ca.* 75% yield.

Nanoparticles were analyzed by NMR *in situ* by detaching the thiols from its surface by treating 500  $\mu$ l of nanoparticles solution with 100  $\mu$ l of a saturated solution of iodine in deuterated solvent.

### 2.2.1. Thiol exchange reaction

**T1-NP-a:** In a sealed vial, under N<sub>2</sub> atmosphere, 0.5 equiv of thiol **1** were added to a MeOH solution (3 mL) of **T2-NP-a** (20 mg) and the solution was reacted at room temperature for 16 hours. The resulting NPs **T1-NP-a** were finally purified by gel permeation chromatography using Sephadex LH-20 resin.

**T1-NP-b:** In a sealed vial, under N<sub>2</sub> atmosphere, 20 equiv of thiol **1** were added to a MeOH solution (2 mL) of **T2-NP-b** (10 mg) and the solution was reacted at room temperature for 16 hours.

The resulting NPs **T1-NP-b** were finally purified by using Amicon<sup>®</sup> Ultra filters of 100 KDa molecular weight cutoff.

### 2.2.2. Characterization of T1-NP-a.

TEM analysis of the different samples of small nanoparticles **T1-NP-a** (Fig. 1) yields an average diameter of 2.1 $\pm$ 0.2 nm. NMR analysis (Fig. 2) indicates monolayer formation (broadening of all

signals), as confirmed by diffusion-filtered experiments (not shown). Composition of T1-NP-a was determined to be 30% after integration of distinctive 1 NMR signals.

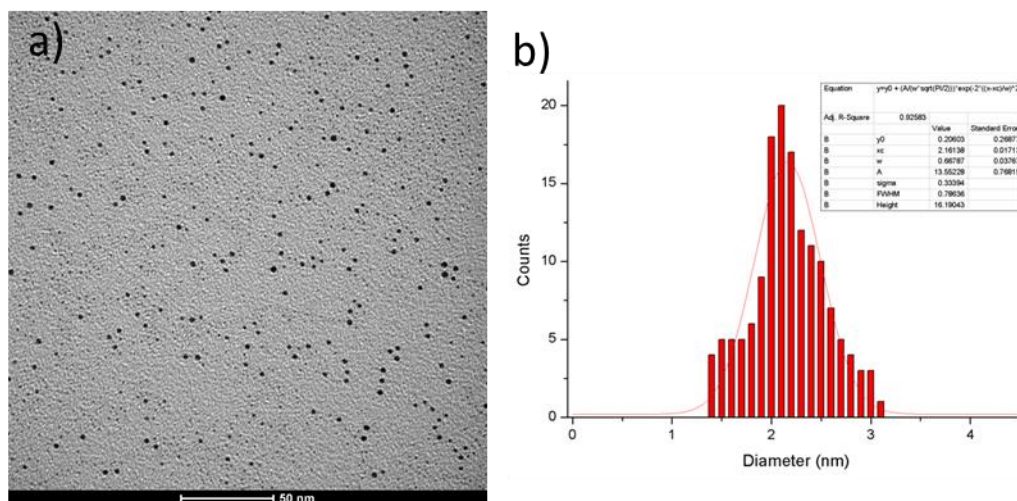


Figure 1. a) Low resolution TEM image of T1-NP-a and b) size distribution (average diameter =  $2.1 \pm 0.2$  nm).

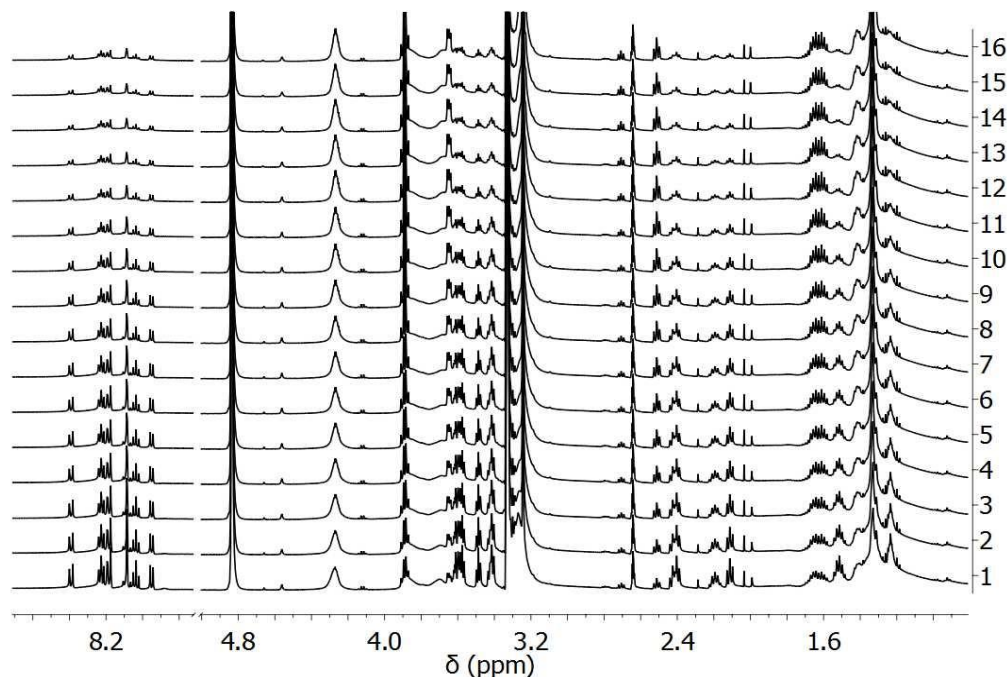
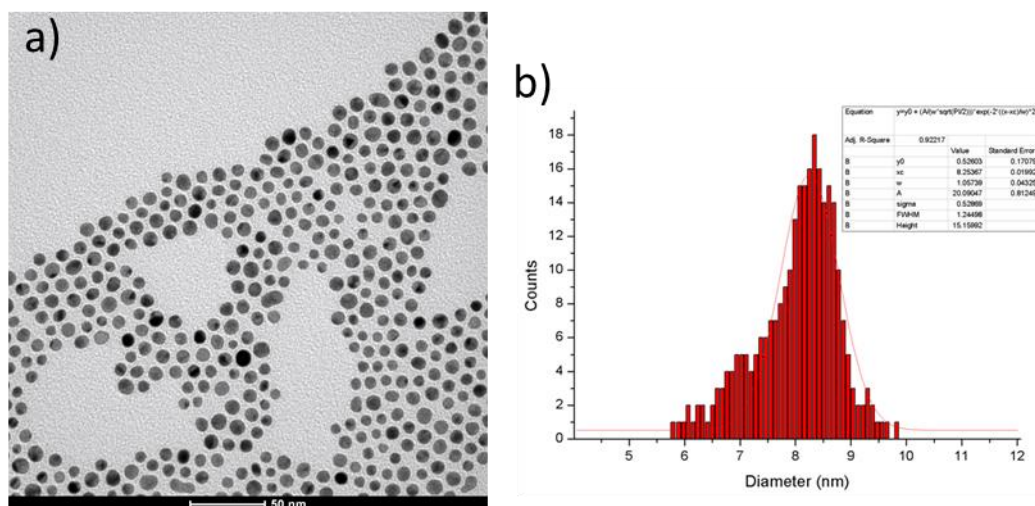


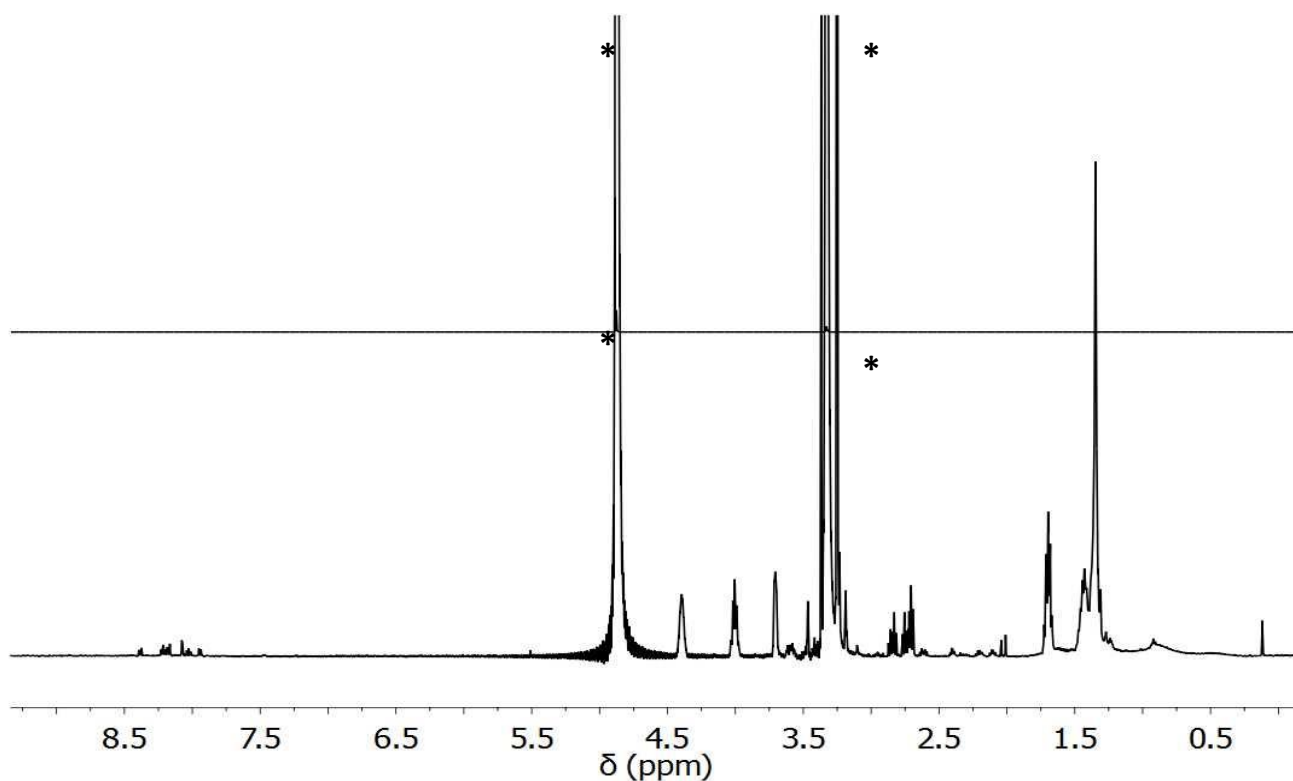
Figure 2.  $^1\text{H}$  NMR spectrum of T1-NP-a thiol exchange reaction. Conditions: 1.2 mg of thiol 1, 3.8 mg of T2-NP-a in 0.6 ml of MeOD. Spectrum 1 was acquired immediately after the addition of thiol 1; subsequent  $^1\text{H}$ -NMR spectra were acquired every hour.

### 2.2.3 Characterization of T1-NP-b.

TEM analysis of the different samples of T1-NP-b (Fig. 3) yields an average diameter of  $8.2 \pm 0.5$  nm. NMR analysis (Fig. 4) indicates monolayer formation (broadening of all signals). Composition of T1-NP-a was determined to be 30% after integration of distinctive 2 NMR signals.



**Figure 3.** a) Low resolution TEM image of **T1-NP-b** and b) size distribution (average diameter  $8.2 \pm 0.5$  nm).



**Figure 4.**  $^1\text{H}$  NMR spectrum of **T1-NP-b** in MeOD before (Top) and after (Bottom)  $\text{I}_2$  oxidation (\* indicates the residual solvents and impurities). No signals are visible in the NP spectrum before oxidation since the slow nanoparticles tumbling rate causes signal broadening beyond detectability.

### 2.3. Decoration of RGO with PyAuNPs

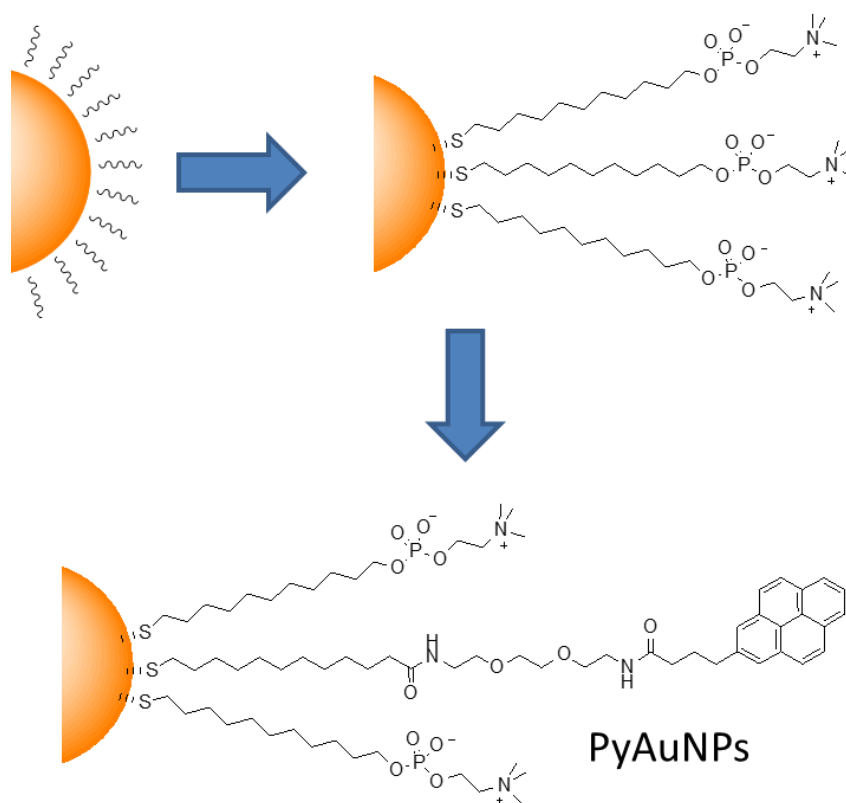
2 mg of **T1-NP-a, b** were dissolved in 5 mL of milliQ  $\text{H}_2\text{O}$ , then 1 mg of RGO (purchased from ACS Material, LLC, product No.: GnP1L-0.5g, and used as received) was added and the suspension was sonicated for 1 hour ( $1.2 \text{ W mL}^{-1}$ ). The resulting solutions turned clear, indicating that the majority of the AuNPs were bound to the graphene and that the formed hybrid material was insoluble in water. The obtained precipitate was purified by washing it in water for 5 times.

## 2.4 Preparation of the SWCNTs-PyAuNPs and SWCNTs-Py hybrids

CoMoCATs nanotubes by SouthWest NanoTechnologies (SWeNT, SG 76, diameter of  $(0.9\pm 0.3)$  nm and high aspect ratio ( $>1000$ )), were dispersed in water at a concentration of  $0.1 \text{ g L}^{-1}$ . 4 mL of the water-SWCNTs mixture were then treated with additional 0.5 mL of, respectively: a) T1-NP-a for the SWCNTs-PyAuNPs sample (T1-NP-a concentration =  $0.4 \text{ mg mL}^{-1}$ ; pyrene concentration in T1-NP-a = 30 wt%) and b) pyrene-derivative for the SWCNTs-Py sample ( $3.5 \times 10^{-5} \text{ M}$ ). The concentrations of the pyrene-derivative and of T1-NP-a solutions were adjusted to ensure the same amount of pyrene in both solutions. To further promote exfoliation of the nanotube bundles and stability of the final suspensions, sodium cholate was further added to the solutions (1 wt%). The two solutions were then sonicated with a tip-sonicator (Bandelin SonoPlus HD 2070) at a power of 60 W and cycles of 0.5 s. After 1h, the suspensions were centrifuged (Hettig Mikro 220R centrifuge) at 3500 rpm at  $27^\circ \text{C}$  for 60 min. The collected supernatant was used for the experiments.

## 3. Results and discussion

Differently from our previous work with SWCNTs,<sup>[32]</sup> this time we choose to test the same 2 nm PyAuNPs and another set of PyAuNPs with a larger diameter (8 nm), thus featuring a plasmonic contribution in the visible range (see the absorption spectrum in Fig. 7). The synthesis of the mixed monolayer particles capped with choline and pyrene derivatives (Fig. 5) was performed according to published procedures.<sup>[32, 33]</sup>



**Figure 5.** Graphical sketch of the synthesis of PyAuNPs.

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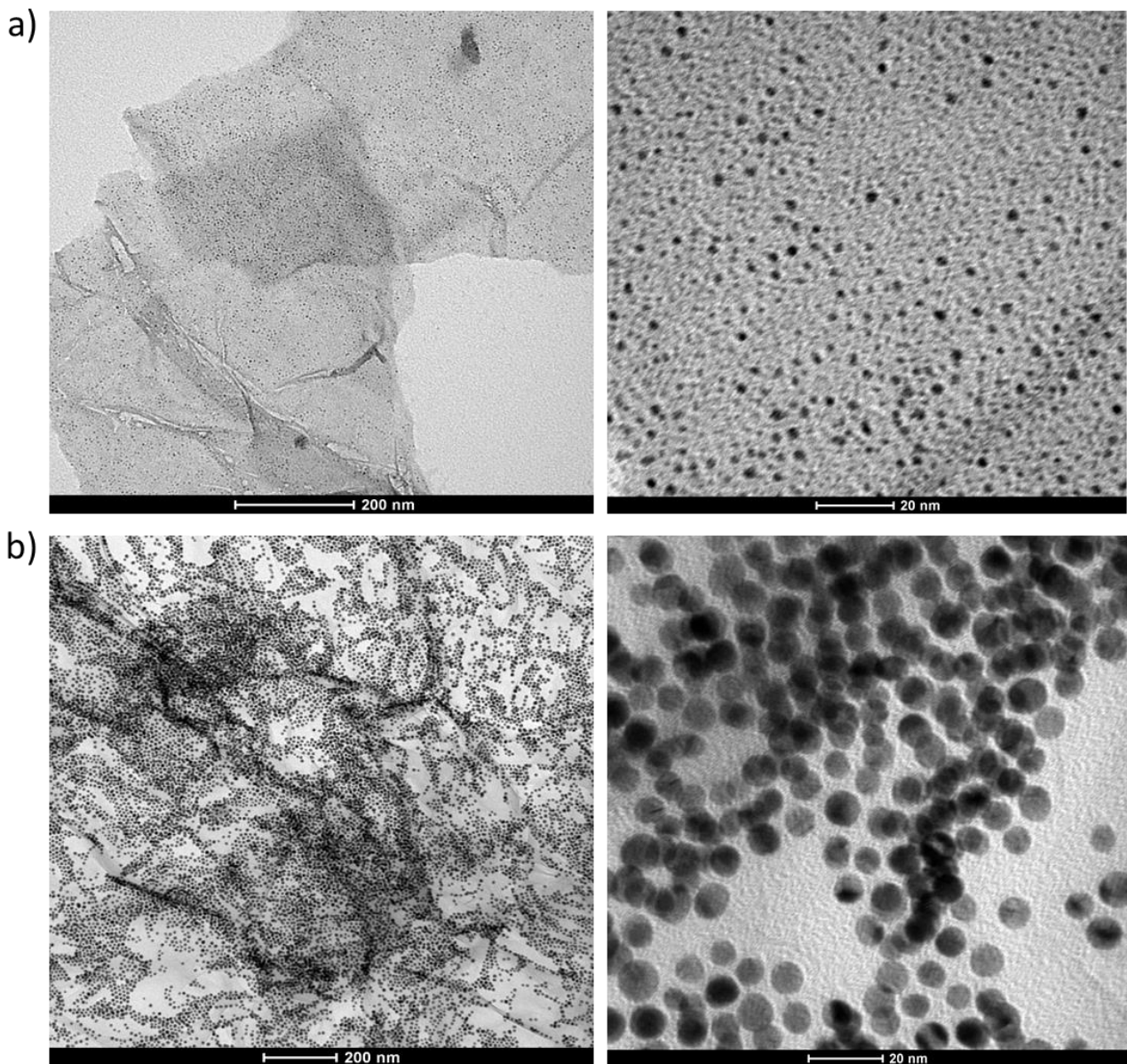
Smaller nanoparticles were prepared by reduction of tetrachloroaurate in toluene with NaBH<sub>4</sub> in the presence of dioctylamine, while larger ones were prepared by silver seeded reduction of tetrachloroaurate in water with sodium citrate. In our previous work,<sup>[32]</sup> we showed that direct formation of the mixed monolayers made by thiols **1** and **2** through simultaneous addition of the two coating molecules did not allow to obtain nanoparticles with **1** content higher than 15%. For this reason, we decided to prepare the mixed monolayers by thiols exchange. **2**-coated AuNPs were prepared by exchanging the weak stabilizing agents (dioctylamine and citrate respectively in the case of 2 and 8 nm diameter nanoparticles) and then subjected to exchange with thiol **1** in methanol (**T1-AuNP-a**). After equilibration (followed by NMR, see Fig. 2) the nanoparticles were purified by size exclusion chromatography and the monolayer composition was determined by <sup>1</sup>H-NMR after thiol detachment by I<sub>2</sub> oxidation.

We found that monolayer coating 2 nm PyAuNPs contained 30% of **1** thiols, while the one coating 8 nm PyAuNp contained 20%. Interestingly, substitution required much larger amounts (20 equivalents) of entering thiols with large nanoparticles than with the small ones (0.5 equivalents), suggesting that the higher curvature favors the mixed monolayer either for kinetic or thermodynamic reasons. Both the nanoparticles are well soluble in water and methanol, and NOE-NMR experiments (not shown) suggest that the pyrene groups are, as expected, bent into the monolayer formed by the alkyl chains of the thiols.

The initial idea in the preparation of the two types nanohybrids of PyAuNPs with RGO was that of obtaining, as it was for the case of the nanohybrids with SWCNTs,<sup>[32]</sup> water solubilization of the GBM by means of the AuNPs bearing zwitterionic pendant choline groups. In this way, multivalency<sup>[31]</sup> on the AuNPs would have been exploited, allowing these last ones to act similarly to surfactants for RGO. Differently from common surfactants anyway, PyAuNPs are expected to be selective for individual sheets rather than for small RGO aggregates, which are sometimes found within standard surfactant micelles dispersing carbon nanomaterials in water. Unfortunately things went differently, being RGO the species driving the solubility profile of the final hybrid. The procedure (described in the Experimental Section) involving addition of pristine RGO to a water solution of the PyAuNPs followed by sonication and ultracentrifugation to eliminate major aggregates, lead to the total precipitation of all the material present in solution, with no residual free particle in the supernatant. This indicates that the affinity of the pyrene functional groups on the AuNPs for RGO is very high and that all the particles are more keen to attach on this last one than to remain in water solution. It is likely that an upper limit to the density of PyAuNPs attached to the RGO surface is given by the lack of residual space on the conjugated carbon based nano-platform. This anyway remains subject of future studies. By looking at the TEM images of the precipitated materials (Fig. 6), it is evident a very

1 high degree of coverage of the RGO surface, characterized by a pronounced homogeneity,  
2 particularly for the case of the PyAuNPs-RGO hybrid with the smaller particles (Fig. 6a). In the  
3 hybrid with the bigger particle, a higher amount of AuNPs aggregates onto the RGO surface is present  
4 (Fig. 6 b).

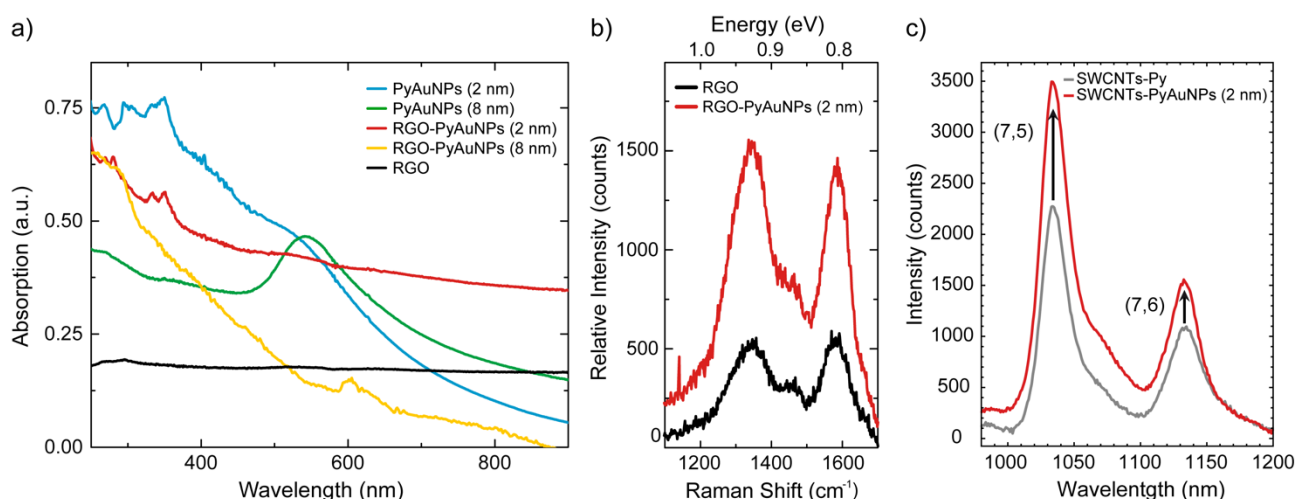
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7 The formation of aggregates of 8 nm PyAuNPs in the hybrid with RGO is also somehow evident from  
8 the UV-visible-NIR spectrum of the hybrid, which features a very weak shoulder over the continuous  
9 absorption contribution of RGO in water (highly scattering and thus completely flat from UV to NIR)  
10 that is red-shifted (very noisy signal with a maximum at around 604 nm) with respect to the plasmonic  
11 band of the free 8 nm PyAuNPs (maximum at 541 nm) and thus can be attributed to aggregated  
12 particles (Fig. 7a).<sup>[36]</sup>  
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61 **Figure 6.** TEM images of PyAuNPs-RGO nanohybrids with a) 2 nm PyAuNPs and b) 8 nm PyAuNPs.

To probe the nature of interaction between PyAuNPs and the RGO surface, the RGO-PyAuNPs composites were analyzed for their Raman spectra. We recorded the spectra of RGO-PyAuNPs (2 nm) composite and compared it to the one of bare RGO (we were not able to get a spectrum from RGO-PyAuNPs -8 nm- composite due to the increased scattering losses by the bigger particle). The comparison between the two spectra is reported in Fig. 7 b. The typical D and G bands of a GBM are found in the Raman spectrum of the RGO-PyAuNPs at 1336 and 1586  $\text{cm}^{-1}$ , respectively. The two main bands for the RGO derivative resemble those present in the Raman spectrum of pristine RGO, but with different  $I_D/I_G$  ratio. This last parameter appears to be slightly higher in the composite, indicating at first approximation the generation of further defects on the RGO surface<sup>[37]</sup> following non-covalent functionalization with PyAuNPs. More properly, the reason for higher  $I_D/I_G$  ratio in the hybrid with respect to the pristine GBM might be related to the reduction in size of single aromatic clusters present on RGO, due to the presence of the particles that act like local interrupters of aromaticity.<sup>[37]</sup> Notwithstanding that, what results to be more striking from the two spectra is the enhancement in the Raman signal of RGO when coupled to AuNPs. The signal of RGO in the hybrid is indeed found to be enhanced by  $\sim 3$  times in the presence of the AuNPs. This observed Raman signal enhancement is similar to what found by Jasuja et al. for composites of AuNPs and graphene oxide<sup>[38]</sup> and paves the way to the application of our hybrid systems as functional platforms for surface enhanced Raman scattering (SERS) detection of analytes.<sup>[39]</sup>

As no direct band-band recombinative emission can be observed from RGO, we tested other semiconducting carbon-based nanosystems to further support the evidence of PyAuNPs-promoted optical response enhancement.



**Figure 7.** a) UV-visible-NIR absorption spectra of PyAuNPs (2 and 8 nm) and their composites with RGO (the absorption of bare RGO is also reported) in water dispersion. b) Raman spectra of RGO compared with RGO-PyAuNPs collected from powder samples. c) Emission spectra of the (7,5) and (7,6) SWCNT species excited at 650 nm in the presence of the Py derivative and of PyAuNPs (2 nm).

1 To this purpose, we realized a novel set of nanohybrids incorporating SWCNTs and analyzed them  
2 through photoluminescence (PL) spectroscopy. Close proximity between gold nanoparticles and  
3 SWCNTs has already been demonstrated to enhance the SWCNT emission yield.<sup>[40, 41]</sup> We  
4 accordingly prepared two set of samples, namely SWCNTs coated just with the pyrene-derivative  
5 (**SWCNTs-Py**) and SWCNTs coated with PyAuNPs (2 nm; **SWCNTs-PyAuNPs**). The samples have  
6 been produced by ensuring that the same amount of SWCNTs was present in both, so that the emission  
7 intensities are directly comparable and related to the emission yields. Fig. 7c compares the emission  
8 spectra collected by the two set of samples, clearly showing the gold-assisted enhancement of the  
9 emission from the hybrid carbon-gold system, interacting by means of the Py functionalities present  
10 on the AuNPs.

#### 11 **4. Conclusions**

12 We have presented here and characterized a new series of RGO-AuNP hybrids obtained through a  
13 non-covalent strategy, in which the assembly of the pristine species proceeds mildly in water at room  
14 temperature. The approach is based on exploiting the versatility of AuNPs for being decorated with  
15 different functional groups, introduced on the surface resorting to Au-S covalent bonds.

16 The selected AuNPs contain indeed a mix monolayer in which pendant pyrene functionalities are  
17 present (PyAuNPs), which are able to establish with the surface of conjugated carbon nanostructures  
18 high-affinity interactions through  $\pi$ - $\pi$  stacking. In this way, a very homogenous and high coverage  
19 of the RGO surface with the AuNPs is achieved, generating well-defined entities which demonstrate  
20 to have the full potential for a use as platforms for SERS analysis. This remains therefore the subject  
21 of our future investigations.

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28 Focus Area NanoScale of the Freie Universität Berlin.

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